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## **10 1 Review And Reinforcement**

10-1 Review and Reinforcement (continued) Find the formula mass of each of the following compounds. Show all your work.

11.  $\text{CaCO}_3$  12.  $\text{MgSO}_4$  13.  $\text{NaOH}$  14.  $\text{KCl}$  15.  $\text{BaF}_2$  o JC 04) 137B  
Find the molar mass of each of the following compounds. Show all your work. 16. nitric acid ( $\text{HNO}_3$ ) (D 17. ammonia ( $\text{NH}_3$ ) 18. calcium sulfate ( $\text{CaSO}_4$ ) 19. potassium sulfide ( $\text{K}_2\text{S}$ )

**Name Date Class 10-1 Review and Reinforcement  
(continued ...**

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Name Date Class 10-1 Review and Reinforcement (continued)

Find the formula mass of each of the following compounds. Show all your work. 11.  $\text{CaCO}_3$  PDF [www.ranchorams.org](http://www.ranchorams.org) '16—1A Review and Reinforcement, pages 9-10 6. Answers Will Vary. When the rates Of forward and reverse reactions are equal, equilibrium is established.

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## **10 1 REVIEW AND REINFORCEMENT CHEMISTRY ANSWERS PDF**

11—1 Review and Reinforcement Stoichiometry Complete the following sentences by 'filling in the appropriate word or phrase from the list below. Each word or phrase maybe used once, more than once, or not at all reactants molar ratio quantitative actual particles. consdrvption of matter coefficients subscripts mass

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10—2 Review and Reinforcement (continued) 9. 0.0621 mol SS = (Q \$ gK2S 10. 65.8gMg= l. atoms Mg Solve each of the following problems as directed. Show all your work. 11. Calculate the number Of atoms in 0.40 mol of sulfur- 12. Calculate the number of atoms in 2.30 mol of silver. G.oux 13. You have a tank of NOZ gas.

## **10-2 Review and Reinforcements**

1. from the given formula, you take the amount of each element, divide by its atomic mass then times 100 2. experimental analysis. a sample of an unknown compound with a mass of 0.2370g is extracted from the roots of a plant. decomposition of

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the sample produces 0.09480 g of C, 0.1264 g of O, and 0.0158 g of H. what is the percent composition ...

### **10-1 and 10-2 Chemistry Flashcards | Quizlet**

10-4 Review and Reinforcement Chemical Measur,cr eat' Define each of the following terms in complete sentences. 1. mole 2. formula mass 3. formula unit 4. Avogadro's number 5. molar mass lithe statement is true. write "true. ' If it is false. change the underlined word or words to make it true. \_\_\_\_ 6.

### **Name Date Ciass of following terms in complete sentences.**

10-1 Review and Reinforcement (continued) Find the formula mass of each of the following compounds Show all your work 11 caC03 12 MgS04 13 NaOH 14 KCl 15 BaF2 o JC 04) 137B Find the molar mass of each of the following compounds Show all

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## **[EPUB] 16 1 Review And Reinforcement Answers Key**

10-1 Review and Reinforcement (continued) Find the formula mass of each of the following compounds Show all your work 11  $\text{CaCO}_3$  12  $\text{MgSO}_4$  13  $\text{NaOH}$  14  $\text{KCl}$  15  $\text{BaF}_2$  o JC 04) 137B Find the molar mass of each of the following compounds Show all your work 16 nitric acid ( $\text{HNO}_3$ ) (D 17 ammonia ( $\text{NH}_3$ ) 18 calcium sulfate ( $\text{CaSO}_4$ ) 19

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16—3 Review and Reinforcement, pages 20-2] 5. endothermic 6. true 7. increases 8. true 9. The Concentrations of reactants and products, pressure, and temperature affect the equilibrium 10. Le Chatelier's principle states that if a change is imposed on a system at equilibrium, the equilibrium position will shift in the



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24—1 Review and Reinforcement (continued) 13. There are 3.29 g of iodine-126 remaining in a sample originally containing 26.3 g of iodine-126. The half-life of iodine-126 is 13 days. How old is the sample? 14. A sample that originally contained 2.5 g of rubidium-87 now contains 1.25 g. The half-life of rubidium-87 is  $6 \times 10^{10}$  years.

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